**Making Ions –** Remember that atoms want a filled outer orbital to be in the most stable state. *Complete the chart below showing what happens for each of the atoms to become an ion*.

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| **Element** | **Lewis Dot** | **# of Valance e-** | **Gain/Lose \_\_\_ e-** | **Valance Charge** |
| Na |  | 1 | L 1 | +1 |
| Be |  |  |  |  |
| Cl |  |  |  |  |
| S |  |  |  |  |
| Al |  |  |  |  |
| Ne |  |  |  |  |
| K |  |  |  |  |
| N |  |  |  |  |
| O |  |  |  |  |
| Ca |  |  |  |  |
| P |  |  |  |  |