Lewis Diagrams and Ionic Bonding Note: Online Textbook

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**Valence Shell –**

Chemical interactions are exchanges of \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, and the only \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ that are free to react with others are the ones in the valence shell.

**Lewis Dot Diagrams:**

Examples:

Nitrogen

Neon

Lithium

Chlorine

Flourine

**Ionic Bonding**

Electron shells have a preference to be \_\_\_\_\_\_\_ or \_\_\_\_\_\_\_\_\_\_\_. In between is considered \_\_\_\_\_\_\_\_\_\_\_.

Lets take a look at the Lewis dot diagram of Lithium and Fluorine.

Observe, one is positively charged, one is negatively charged. They will attract. Then they are “ionically bonded”

**Ionic Bond:**

Examples

Magnesium and Oxygen

More Examples:

Sodium and Chlorine

Calcium and Flourine

Potassium and Nitrogen