**6.2 Reaction Rates**

In order for a chemical reaction to take place, the reactants must \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

The collision between the molecules in a chemical reaction provides the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ needed to break the necessary bonds so that new bonds can be formed.

Sometimes, even if there is a collision, not enough kinetic energy is available to be transferred — the molecules aren’t moving fast enough.

By changing how the reactants collide, it is possible to control how quickly or slowly a reaction will occur.

**Factors affecting Reaction Rate**

**Temperature**

**Concentration**

**Surface Area**

**Catalyst and Inhibitors**



\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ are substances that make things easier for reacting. One way or another, it reduces the energy required for a reaction to happen. It is an intermediary in the reaction, but isn’t used as part of the reactants nor becomes part of the products.

Inhibitors are substances that, like catalysts, are not used as reactants or products, but unlike catalysts increase the energy required to make reactions happen.