**Acids, Bases, and Salts**

**Acid-Base Reactions**

pH

$$HCl + NaOH \rightarrow \\_\\_\\_\\_\\_\\_\\_\\_\\_ + \\_\\_\\_\\_\\_\\_\\_\\_\\_$$

 Type

When an acid and a base are combined

* It produces
* It is called

Examples:

$$2CH\_{3}COOH+Ca\left(OH\right)\_{2}\rightarrow 2H\_{2}O+Ca\left(CH\_{3}COO\right)\_{2}$$

$$H\_{2}SO\_{4}+2NH\_{4}OH\rightarrow 2H\_{2}O+(NH\_{4})\_{2}SO\_{4}$$

**Predicting Results**

Given that a neutralization reaction always produce water and a salt. Predict the products in the reactions:

$$NaOH + H\_{2}CO\_{3} \rightarrow $$

$$HNO\_{3}+KOH\rightarrow $$

$$H\_{3}PO\_{4}+ Al\left(OH\right)\_{3}\rightarrow $$

**Identifying and Naming Acids, Bases, and Salts**

Acids

* Identification:
* Naming:
	+
	+
	+

Bases

* Identification:
* Naming:

Salts

* Identification:
* Naming: