**Polyatomic Ions**

Some covalent and ionic compounds occur very regularly and may gain or lose some extra electrons to become \_\_\_\_\_\_\_\_. When this happens a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ becomes an ion. We call these \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ ions.

Many polyatomic ions are necessary parts of many things in the world. One polyatomic ion, **carbonate** \_\_\_\_\_\_\_\_\_ helps form the shells of bird eggs and also is a primary component on the enamel of your teeth. A compound made from two polyatomic ions: **ammonium** \_\_\_\_\_\_\_\_ and **nitrate** \_\_\_\_\_\_\_\_\_ is one of the most important fertilizers in the world.

**Using Polyatomic Ions in Formulae**

The rules for making compounds using polyatomic ions are very similar to how we make compound without polyatomic ions.

A couple additional rules:

* The **subscripts** (lower letters) on a polyatomic ion are \_\_\_\_\_ allowed to be \_\_\_\_\_\_\_\_\_\_\_
* When more than one polyatomic ion is needed to balance the charge, put a pair of \_\_\_\_\_\_\_\_\_\_\_\_ around the **whole** polyatomic ion

**Examples:**

Potassium and Dichromate Iron (III) and Nitrate

Strontium and Sulfate Ammonium and Phosphate